

IONIC COMPOUNDS: Names and Formulas

1. Write the formulas for the following compounds.

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|------------------------|------------------------------------|----------------------------|------------------------------------|
| a. magnesium oxide | <u>MgO</u> | k. tin (II) fluoride | <u>SnF₂</u> |
| b. aluminum nitride | <u>AlN</u> | l. lead (IV) nitride | <u>Pb₃N₄</u> |
| c. potassium sulfide | <u>K₂S</u> | m. iron (III) chloride | <u>FeCl₃</u> |
| d. calcium bromide | <u>CaBr₂</u> | n. copper (I) oxide | <u>Ca₂O</u> |
| e. aluminum sulfide | <u>Al₂S₃</u> | o. antimony (III) sulfide | <u>Sb₂S₃</u> |
| f. beryllium oxide | <u>BeO</u> | p. mercury (II) oxide | <u>HgO</u> |
| g. strontium phosphide | <u>Sr₃P₂</u> | q. tin (IV) iodide | <u>SnI₄</u> |
| h. sodium fluoride | <u>NaF</u> | r. arsenic (III) phosphide | <u>AsP</u> |
| i. lithium selenide | <u>Li₂Se</u> | s. cobalt (II) sulphide | <u>CoS</u> |
| j. barium oxide | <u>BaO</u> | t. tin (IV) sulphide | <u>SnS₂</u> |

2. Write the names for the following compounds.

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|-----------------------------------|----------------------------|----------------------|--------------------------------|
| a. Li ₂ O | <u>lithium oxide</u> | k. PbS | <u>lead (II) sulphide</u> |
| b. AlCl ₃ | <u>aluminium chloride</u> | l. SnO ₂ | <u>tin (IV) oxide</u> |
| c. MgS | <u>magnesium sulphide</u> | m. NiO | <u>nickel (II) oxide</u> |
| d. CaF ₂ | <u>calcium fluoride</u> | n. CuI ₂ | <u>copper (II) iodide</u> |
| e. Al ₂ O ₃ | <u>aluminum oxide</u> | o. PbCl ₄ | <u>lead (IV) chloride</u> |
| f. BeF ₂ | <u>beryllium fluoride</u> | p. FeP | <u>iron (III) phosphide</u> |
| g. K ₃ P | <u>potassium phosphide</u> | q. AuBr ₃ | <u>gold (III) bromide</u> |
| h. Mg ₃ P ₂ | <u>magnesium phosphide</u> | r. Hg ₂ S | <u>mercury (II) sulphide</u> |
| i. CaO | <u>calcium oxide</u> | s. SbF ₃ | <u>antimony (III) fluoride</u> |
| j. Ag ₂ S | <u>silver sulphide</u> | t. MnO ₂ | <u>manganese (IV) oxide</u> |